

## 6.3 Review Worksheet

Write the ions for the following

Carbonate \_\_\_\_\_

Phosphite \_\_\_\_\_

Fluoride \_\_\_\_\_

Iron (II) \_\_\_\_\_

Phosphate \_\_\_\_\_

Copper (I) \_\_\_\_\_

Titanium (III) \_\_\_\_\_

Beryllium \_\_\_\_\_

Aluminum \_\_\_\_\_

Name the following ions

$\text{Ca}^{+2}$  \_\_\_\_\_

$\text{Cl}^-$  \_\_\_\_\_

$\text{HCO}_3^-$  \_\_\_\_\_

$\text{Cu}^{+2}$  \_\_\_\_\_

$\text{N}^{-3}$  \_\_\_\_\_

Na \_\_\_\_\_

O \_\_\_\_\_

Sr \_\_\_\_\_

$\text{Cr}^{+2}$  \_\_\_\_\_

$\text{NO}_2$  \_\_\_\_\_

$\text{SO}_3$  \_\_\_\_\_

Name the following compounds

$\text{AlF}_3$  \_\_\_\_\_

$\text{Ba}(\text{CN})_2$  \_\_\_\_\_

$\text{BeF}_2$  \_\_\_\_\_

$\text{BeO}$  \_\_\_\_\_

$\text{CBr}_4$  \_\_\_\_\_

$\text{CO}$  \_\_\_\_\_

$\text{CuO}$  \_\_\_\_\_

$\text{FeO}$  \_\_\_\_\_

$\text{KClO}_2$  \_\_\_\_\_

$\text{KClO}_3$  \_\_\_\_\_

$\text{KNO}_2$  \_\_\_\_\_

$\text{LiF}$  \_\_\_\_\_

$\text{MgO}$  \_\_\_\_\_

$\text{MgSO}_4$  \_\_\_\_\_

$\text{N}_2\text{O}_3$  \_\_\_\_\_

$\text{Na}_2\text{O}$  \_\_\_\_\_

$\text{NaBr}$  \_\_\_\_\_

$\text{NaF}$  \_\_\_\_\_

$\text{NF}_3$  \_\_\_\_\_

$\text{Ni}_2\text{O}_3$  \_\_\_\_\_

$\text{TiO}_2$  \_\_\_\_\_

$\text{Mn}(\text{CO}_3)_3$  \_\_\_\_\_

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Write the chemical formula for the following

Copper(II) oxide \_\_\_\_\_

Sodium Nitride \_\_\_\_\_

Magnesium Chloride \_\_\_\_\_

Carbon tetrachloride \_\_\_\_\_

Dinitrogen tetroxide \_\_\_\_\_

Potassium sulfate \_\_\_\_\_

Barium fluoride \_\_\_\_\_

Calcium acetate \_\_\_\_\_

Rubidium sulfate \_\_\_\_\_

Sodium nitrite \_\_\_\_\_

Silver (III) oxide \_\_\_\_\_

Cadmium (I) oxide \_\_\_\_\_

Cobalt (III) hydroxide \_\_\_\_\_

Magnesium chloride \_\_\_\_\_

Barium sulfide \_\_\_\_\_

Lithium oxide \_\_\_\_\_

Potassium hydroxide \_\_\_\_\_

Disulfur nanabromide \_\_\_\_\_

Iron (IV) Carbonate \_\_\_\_\_

Mercury (I) Nitride \_\_\_\_\_